

of the physicists and it seems quite certain that at some future time our theories of chemical combination must be brought into harmony with the known facts about spectral lines, absorption bands, color and other phenomena of light and radiation.

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WILLIAM A. NOYES

A Method for Reproducing Graphs in Quantity.—Graphs, etc., are plotted with waterproof india ink on the ordinary graph (or coördinate) paper. This is dipped in a saturated solution of purified, colorless paraffin oil (such as Stanolind-Liquid Paraffin) in chloroform and allowed to stand for $\frac{1}{2}$ to 1 minute. It is then removed from the solution and allowed to drain as long as convenient. The excess of oil is wiped off and the sheet allowed to dry in the air until ready to use.

Using paper thus prepared, as a negative, photographic copies can be made on any of the bromide or blue print papers by direct printing. The bromide reproductions are particularly satisfactory and resemble photo-stats.

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[CONTRIBUTION FROM THE KENT CHEMICAL LABORATORY OF THE UNIVERSITY OF CHICAGO]

THE BEHAVIOR OF MERCURIC SALTS OF ORGANIC ACIDS TOWARD HEAT¹

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In a previous communication Kharasch⁴ called attention to the fact that mercuric salts of certain aromatic carboxylic acids when heated split off carbon dioxide and the mercury becomes attached to the carbon originally bound to the carboxyl group. That reaction is applicable only to those aromatic carboxylic acids which split off carbon dioxide at their respective melting points or at slightly higher temperatures. In the case of the mercury salts of aromatic carboxylic acids, which do not split off carbon dioxide

¹ Read before the Organic Division of the American Chemical Society at the Pittsburgh Meeting, September, 1922.

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³ The material presented here is used by F. W. Staveley in his dissertation presented in partial fulfilment of the requirements for the Degree of Doctor of Philosophy at the University of Chicago.

⁴ Kharasch, *THIS JOURNAL*, **43**, 2238 (1921).